

To make an electrochemical cell, you need:

ELECTRODES:

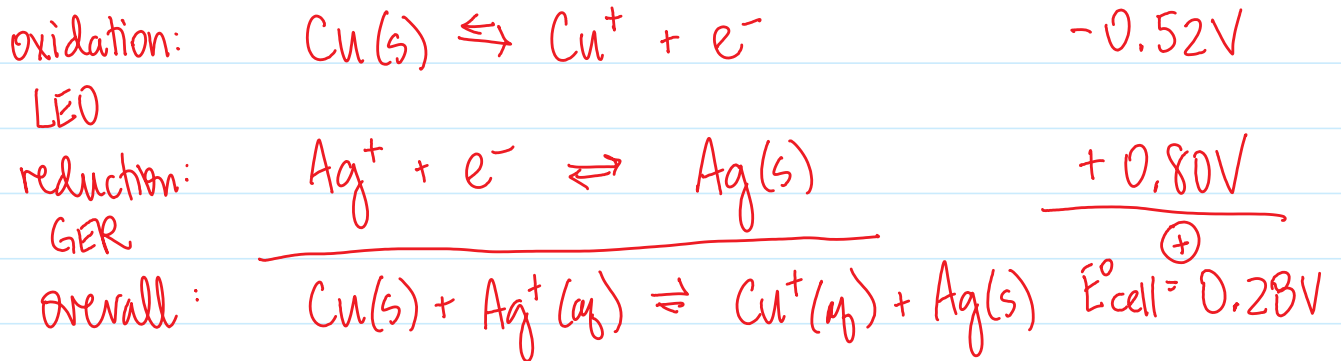
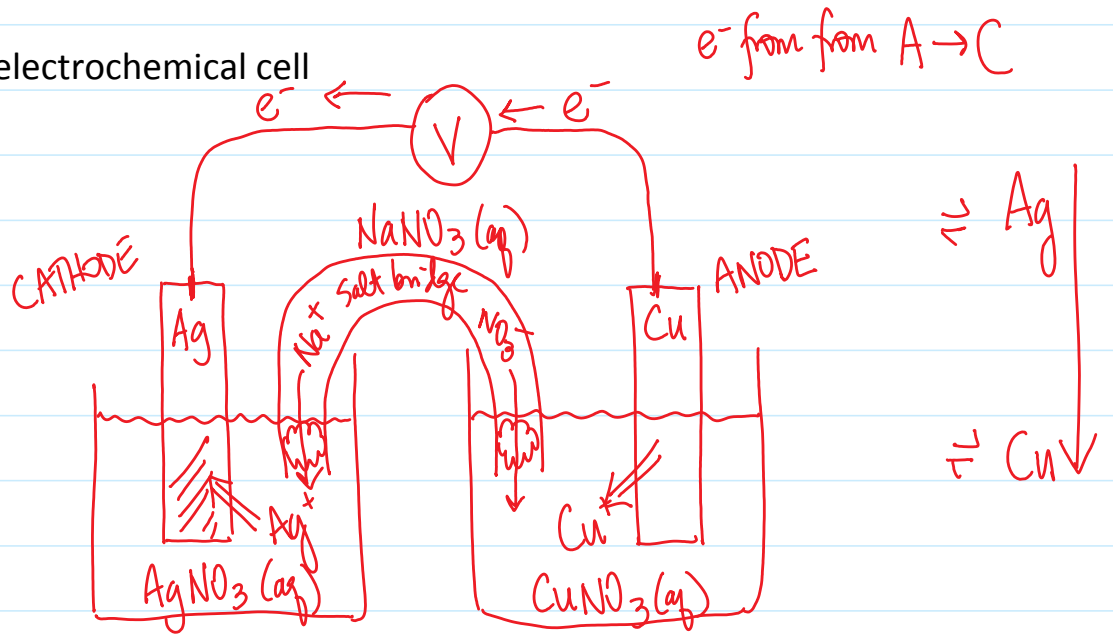
- a. ANODE - where oxidation occurs, ⊖
- b. CATHODE - where reduction occurs, ⊕

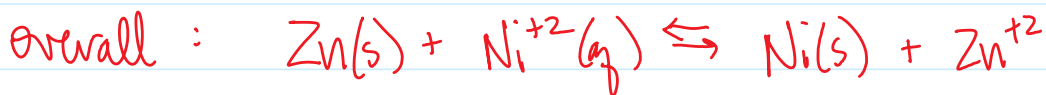
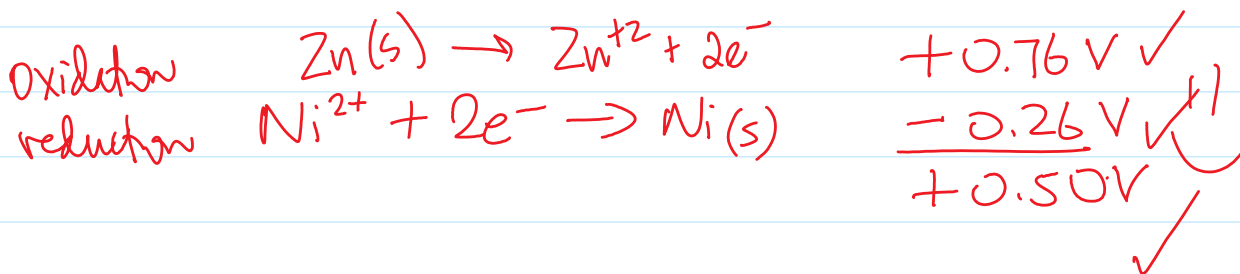
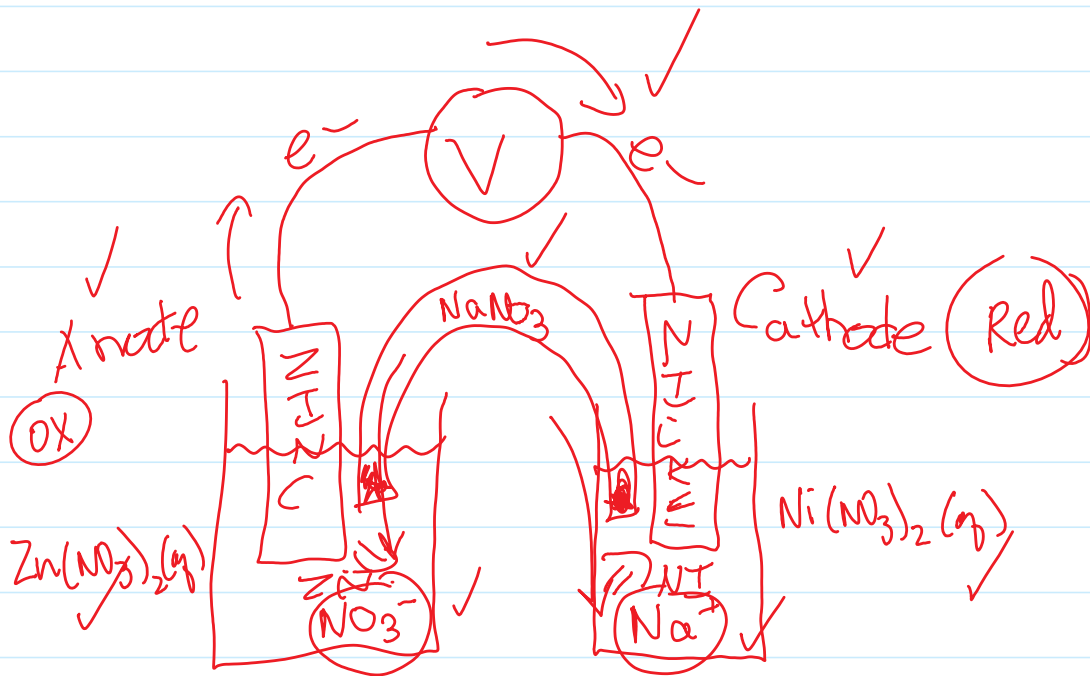
MEMORY HELP: An OX CARE D

anode oxidation cathode reduction

How does it work?

Ex. Ag and Cu electrochemical cell





if you add 3.0M $Ni(NO_3)_2$

V ↑

why?

incr $[Ni^{2+}]$

equil shifts to the products

add 3.0M $Zn(NO_3)_2$

$[Zn^{2+}] \uparrow$

shifts to the reactants

∴ V ↓